

Chapter 3

Section 1

1. Solid- has a definite shape and a definite volume.
2. Crystalline solids- solids that are made up of crystals.
3. Amorphous solids- the particles are not arranged in a regular pattern.
4. Liquid- has a definite volume but no shape of its own.
5. Fluid- meaning of a “substance that flows.”
6. Surface tension- is the result of an inward pull among the molecules of a liquid that brings the molecules on the surface together.
7. Viscosity- a liquid’s resistance to flowing.
8. Gas- can change volume very easily.

Section 2

1. Melting- the change in state from a solid to a liquid.
2. Melting point- in most pure substances, melting occurs at a characteristic temperature.
3. Freezing- the change from a liquid to a solid.
4. Vaporization- the change from a liquid to a gas.
5. Evaporation- Vaporization that takes place only the surface of a liquid.
6. Boiling- occurs when a liquid changes to a gas below its surface as well as at the surface.
7. Boiling point- the temperature at which a liquid boils.
8. Condensation- is the change in state from a gas to a liquid.
9. Sublimation- occurs when the surface particles of a solid gain enough energy that they form a gas.

Section 3

1. Pressure- of the gas is the force of its outward push decided by the area of the walls of the container.
2. Directly proportional- when a graph of two variables is a straight line passing through the origin.
3. Inversely proportional- when the product of two variables is a constant.